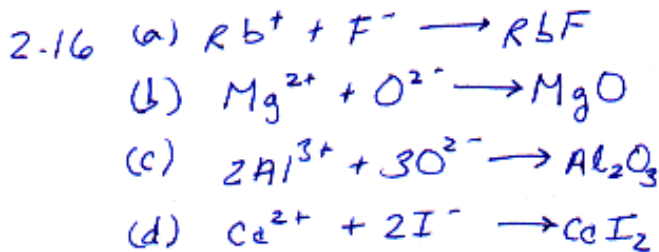


2.15 ~~compounds~~ ionic compounds form by electrostatic interactions. Those atoms that tend to lose e⁻ (metals, those elements on the left to middle of the periodic table) form ionic compounds with those that tend to gain e⁻ (nonmetals, those elements on the right of the periodic table).

- (a) ionic
 (b) both Na & Ca tend to lose e⁻. they do not react together
 (c) N & O form covalent molecules like N₂O
 (d) C & H form covalent molecules like CH₄



2.20

SALT	% (mass) Cl	% (mole) Cl
NaCl	$\frac{35.453 \text{ g Cl}}{58.443 \text{ g NaCl}} \times 100\% = 60.663\%$	$\frac{1 \text{ mol Cl}}{1 \text{ mol Na} + 1 \text{ mol Cl}} \times 100\% = 50\%$
KCl	$\frac{35.453 \text{ g Cl}}{74.55 \text{ g KCl}} \times 100\% = 47.555\%$	$\frac{1 \text{ mole Cl}}{1 \text{ mole K} + 1 \text{ mole Cl}} \times 100\% = 50\%$
RbCl	$\frac{35.453 \text{ g Cl}}{120.921 \text{ g RbCl}} \times 100\% = 29.279\%$	$\frac{1 \text{ mole Cl}}{1 \text{ mol Rb} + 1 \text{ mol Cl}} \times 100\% = 50\%$