


PROBLEMS

 Problems marked with an icon may be more conveniently solved on a personal computer with a mathematical program.

13.1 Since the energy of a molecular quantum state is divided by kT in the Boltzmann distribution, it is of interest to calculate the temperature at which kT is equal to the energy of photons of different wavelengths. Calculate the temperature at which kT is equal to the energy of photons of wavelength 10^3 cm, 10^{-1} cm, 10^{-3} cm, and 10^{-5} cm.

13.2 Most chemical reactions require activation energies ranging between 40 and 400 kJ mol⁻¹. What are the equivalents of 40 and 400 kJ mol⁻¹ in terms of (a) nm, (b) wavenumbers, and (c) electron volts?

13.3 (a) What vibrational frequency in wavenumbers corresponds to a thermal energy of kT at 25 °C? (b) What is the wavelength of this radiation?

13.4 Show that equation 13.17 is a solution of equation 13.9 by differentiating equation 13.17 and substituting it into equation 13.9.

5-7
13.5 Calculate the reduced mass and the moment of inertia of D³⁵Cl, given that $R_e = 127.5$ pm.

13.6 The H—O—H bond angle for ¹H₂O is 104.5°, and the H—O bond length is 95.72 pm. What is the moment of inertia of H₂O about its C₂ axis?

13.7 Some of the following gas molecules have pure microwave absorption spectra and some do not: N₂, HBr, CCl₄, CH₃CH₃, CH₃CH₂OH, H₂O, CO₂, O₂. What is the gross selection rule for rotational spectra, and which molecules satisfy it?

5-7
13.8 Calculate the frequency in wavenumbers and the wavelength in cm of the first rotational transition ($J = 0 \rightarrow 1$) for D³⁵Cl.

5-7
13.9 The pure rotational spectrum of ¹²C¹⁶O has transitions at 3.863 and 7.725 cm⁻¹. Calculate the internuclear distance in ¹²C¹⁶O. Predict the positions, in cm⁻¹, of the next two lines.

5-7
13.10 Assume the bond distances in ¹³C¹⁶O, ¹³C¹⁷O, and ¹²C¹⁷O are the same as in ¹²C¹⁶O. Calculate the position, in