Introduction to Natural Science, Fall 2006
Chemistry Workshop – Week 4

1. Determine the ions formed by the following elements. Write their formulas, names, and determine the number of protons and electrons in each of the ions.
    - rubidium (Rb)
    - strontium (Sr)
    - phosphorus (P)
    - tin (Sn)
    - bromine (Br)
    - gallium (Ga)
    - cadmium (Cd)

2. Complete the following table by writing the name or formula of the compound. If the compound is ionic, write the anions and cations. If the compound is molecular, write "molecular". The first two are done for you.

name	**formula**	**If ionic, write the ions, otherwise write "molecular"**
sulfur difluoride	SF_2	molecular
sodium oxide	Na_2O	Na^+ and O^{2-}
diboron trioxide		
sulfur hexafluoride		
	CaS	
	$TiCl_4$	
carbon tetrachloride		
strontium bromide		
	Rb_2O	
cobalt(II) sulfide		
aluminum carbonate		
	$FeBr_3$	
copper(II) nitrate		
	AgCl	
potassium hydrogen sulfide		
manganese(IV) sulfide		
magnesium phosphide		
ammonium nitride		
	$Pb(NO_3)_2$	
diphophorus pentoxide		
	CrP	
sulfur dioxide		
iron(III) nitrate		
	$Al(OH)_3$	
dinitrogen tetroxide		
	Hg_2O	
cesium perchlorate		