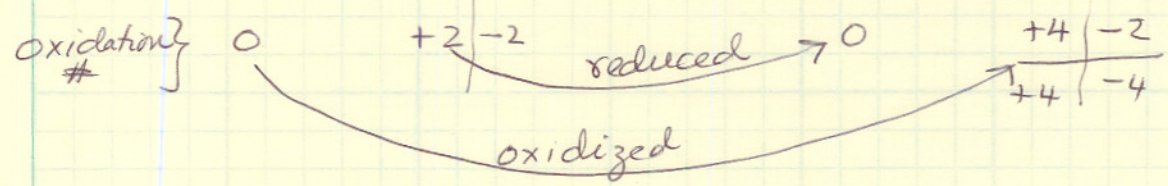
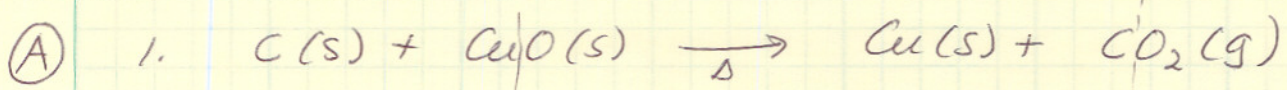
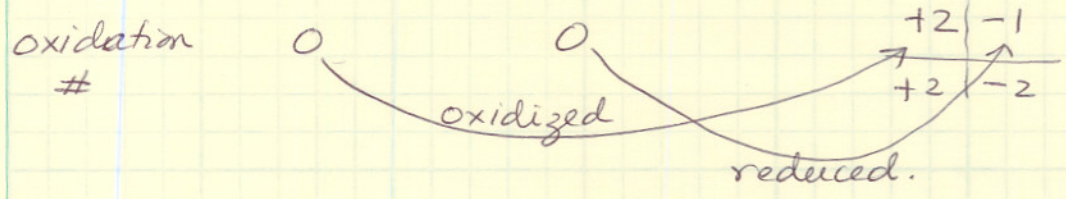
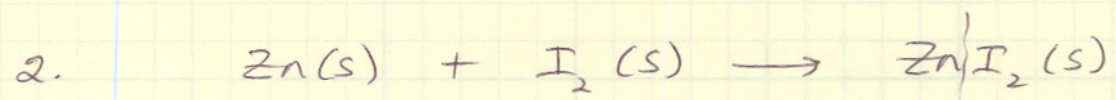


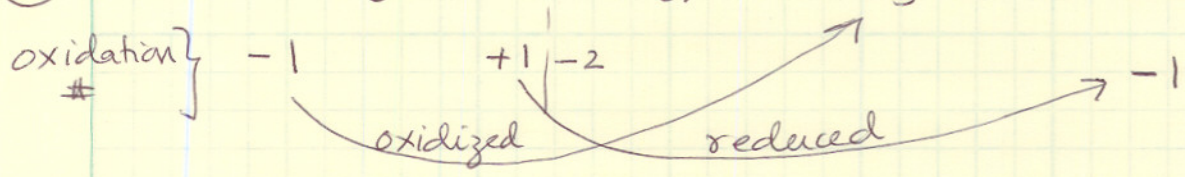
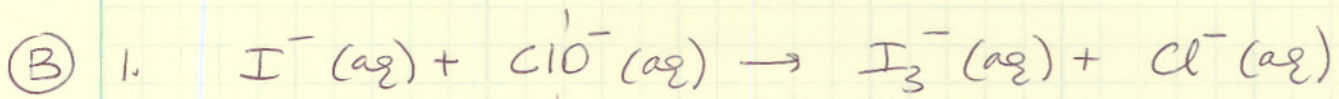
ANSWERS TO REDOX REACTIONS



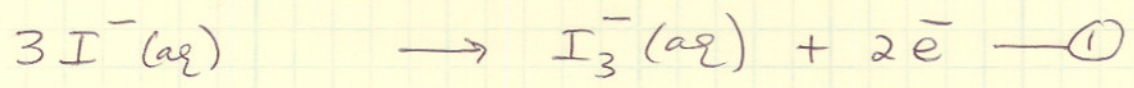
C is oxidized.
 Reducing agent is C(s).
 Cu is reduced. CuO is the oxidizing agent.



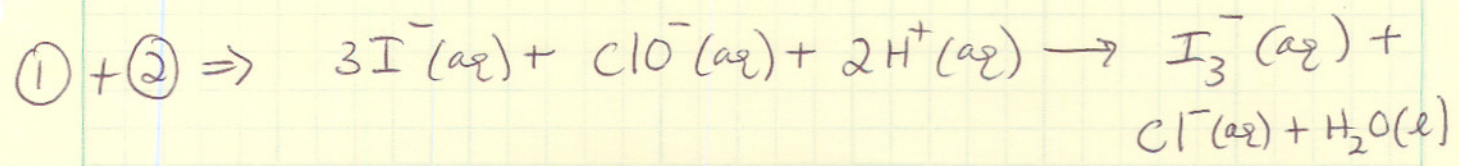
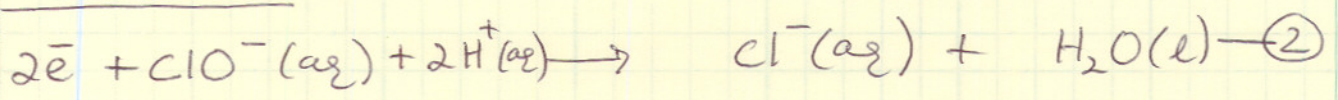
Zn is oxidized. Zn is the reducing agent.
 I is reduced. I₂ is the oxidizing agent.



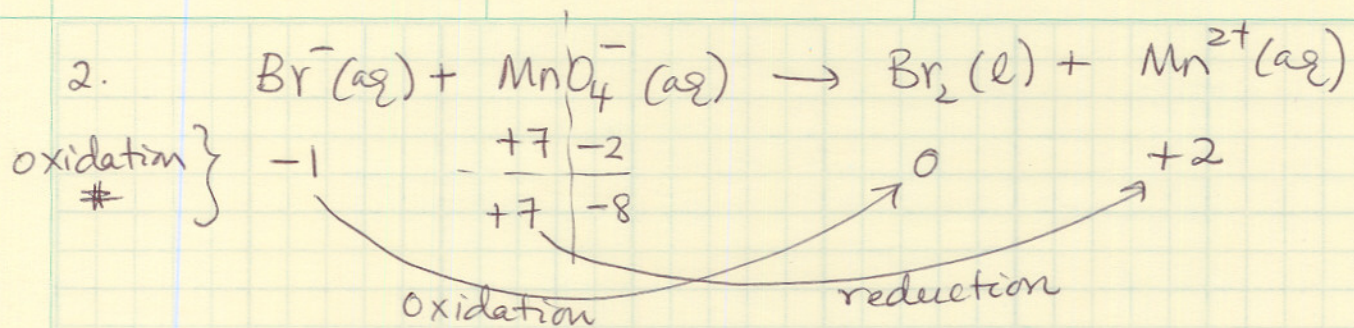
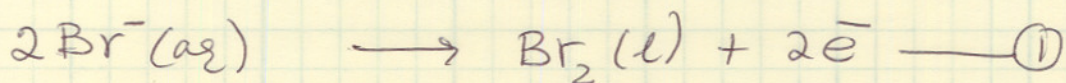
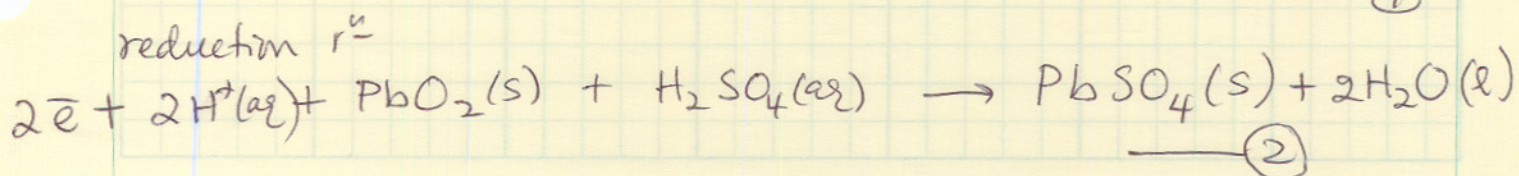
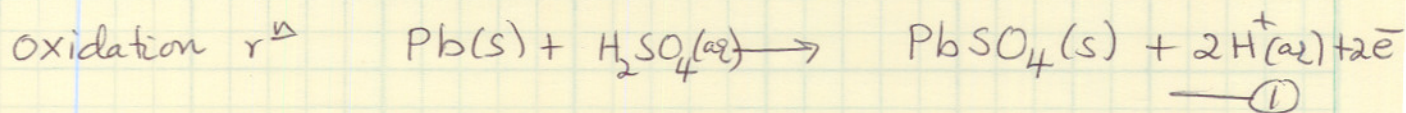
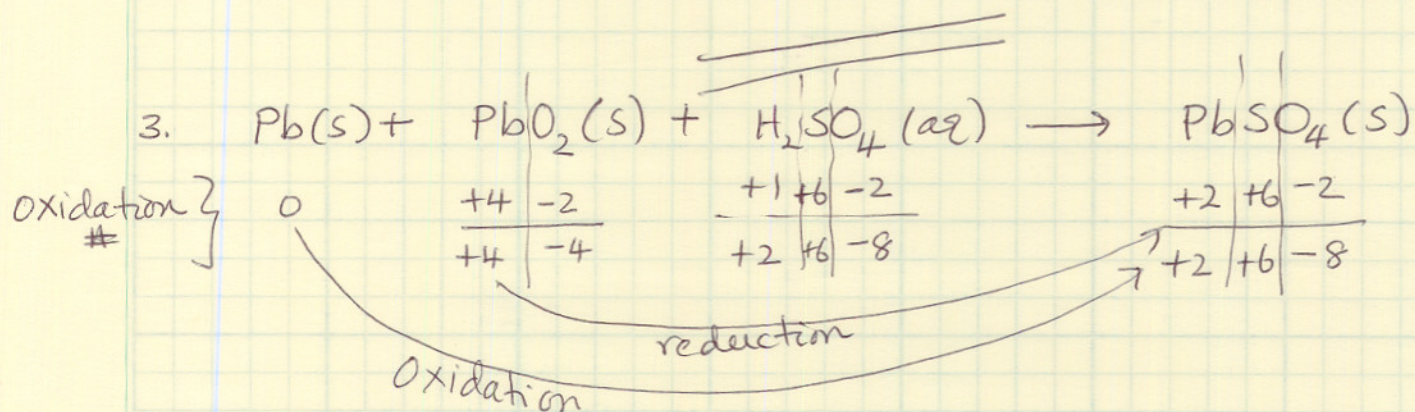
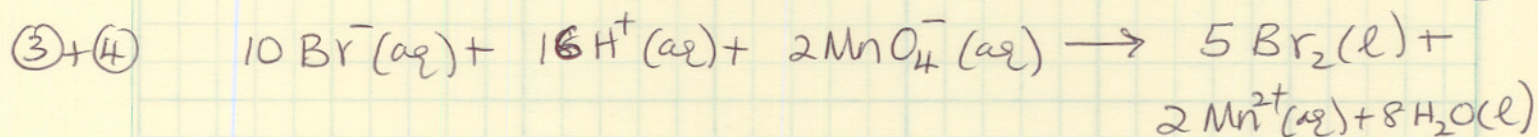
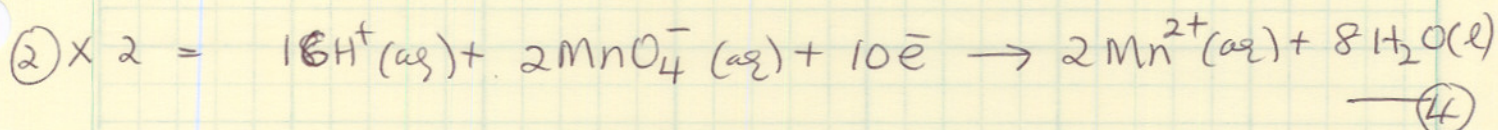
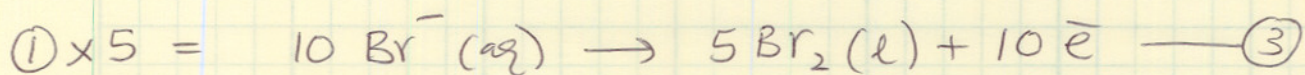
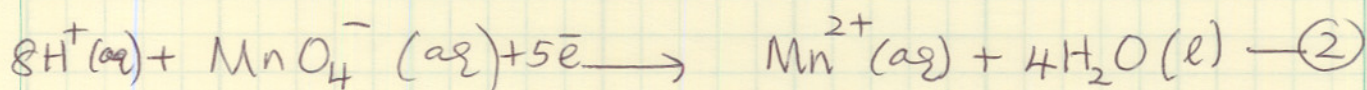
oxidation rⁿ



reduction rⁿ

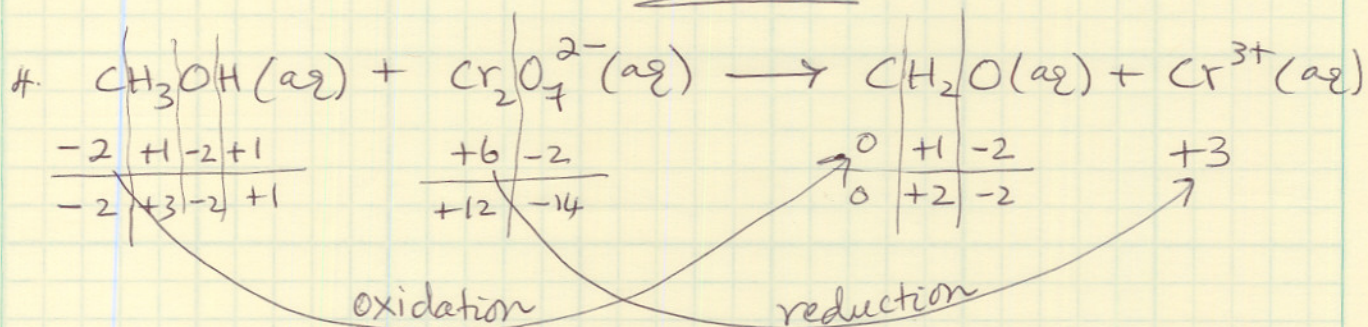
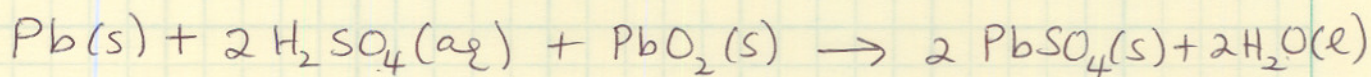
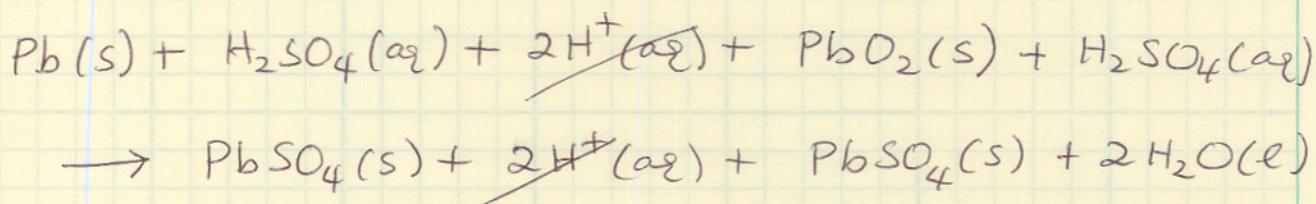


2/

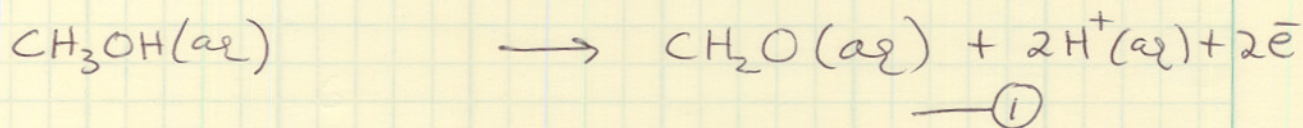
oxidation rⁿreduction rⁿ

③

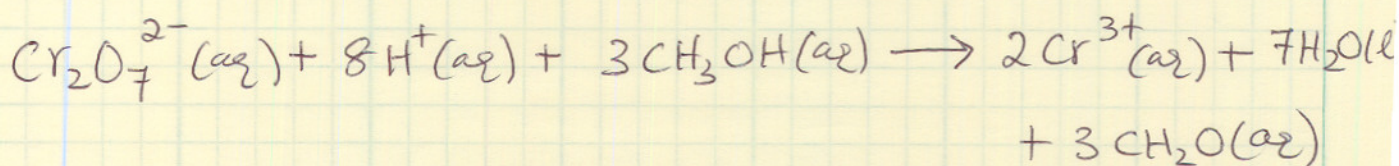
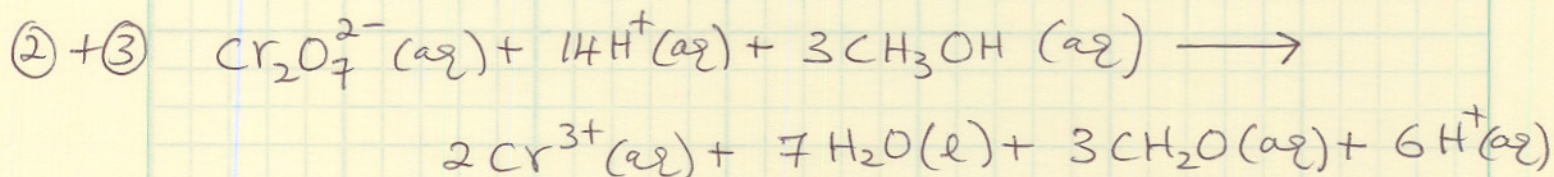
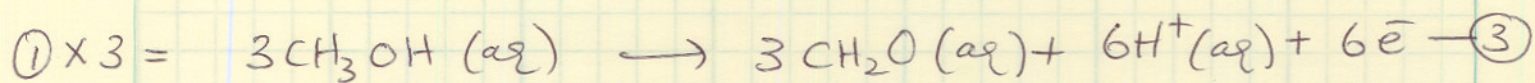
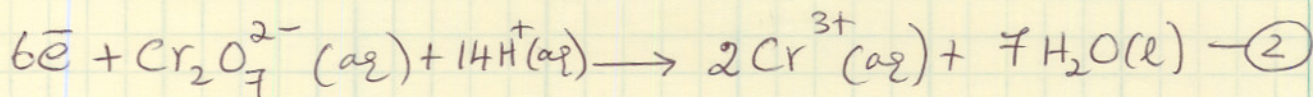
① + ②

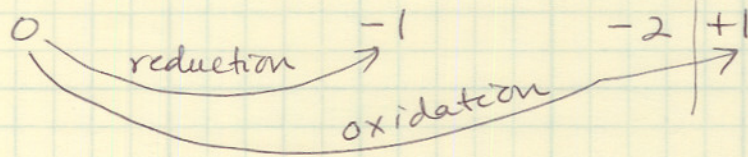
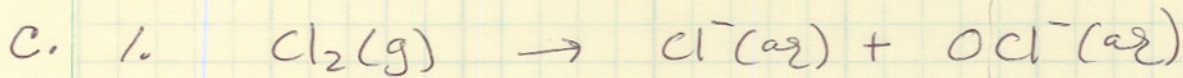


oxidation rⁿ

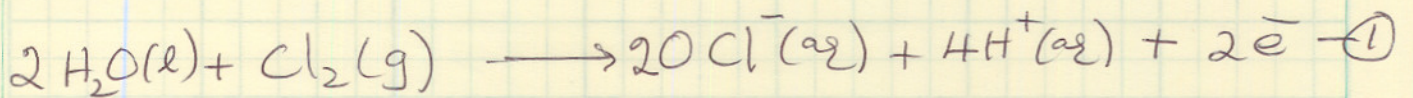


reduction rⁿ

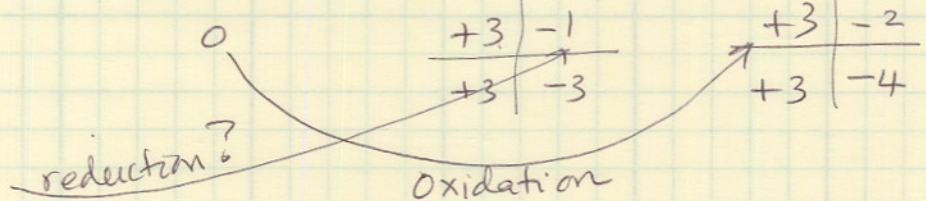
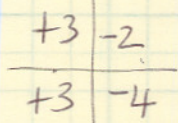
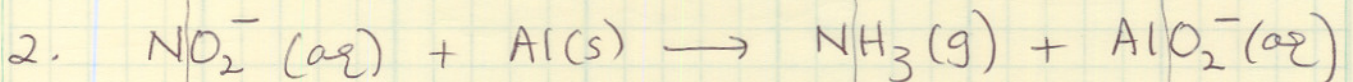
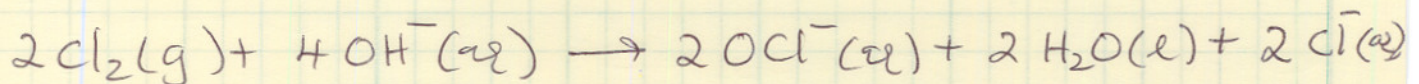
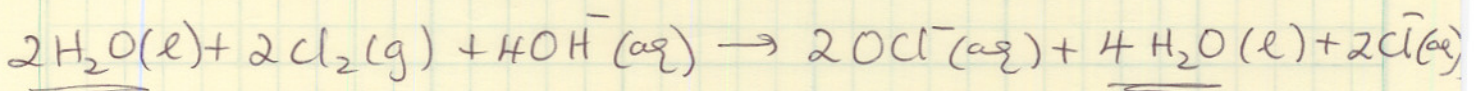
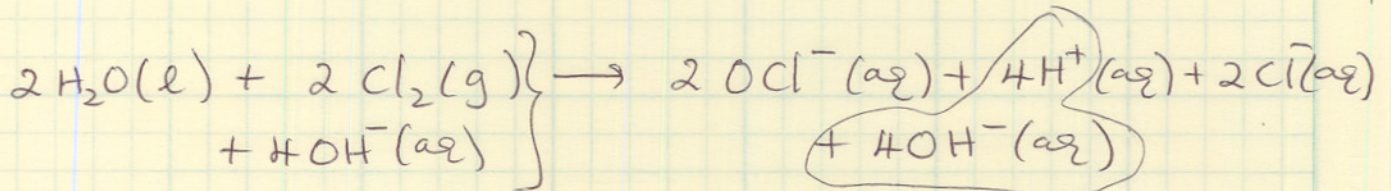
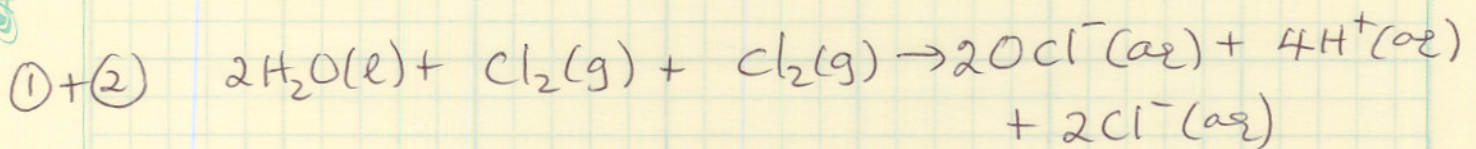
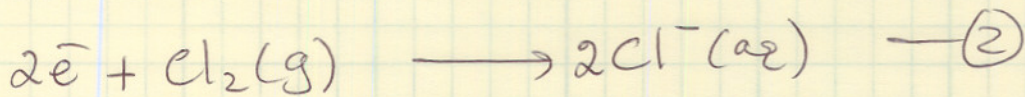




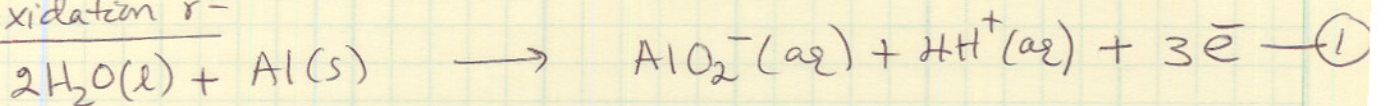
oxidation rⁿ



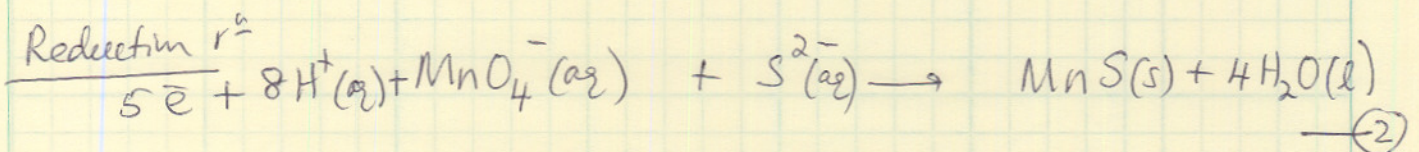
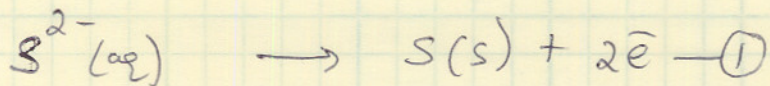
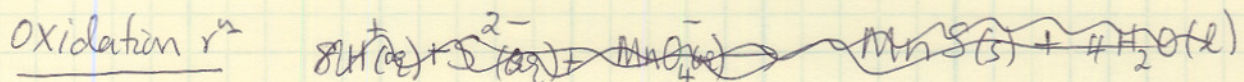
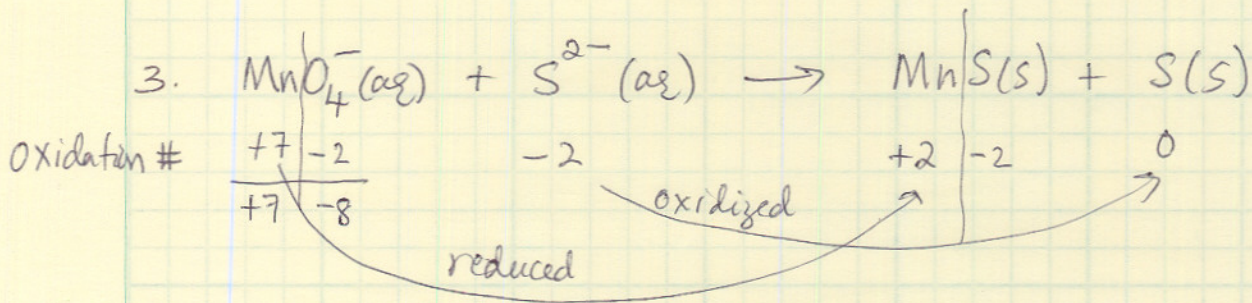
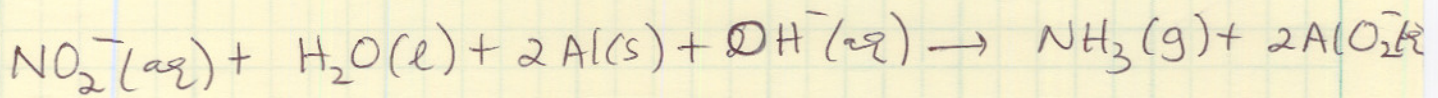
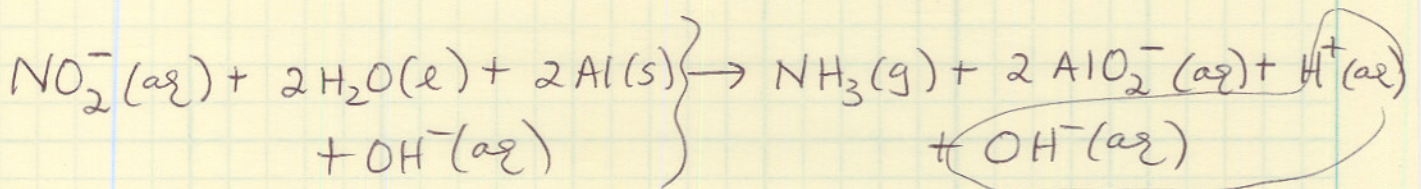
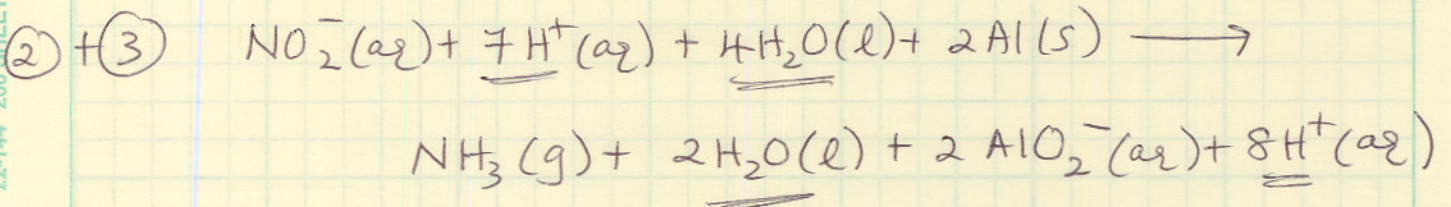
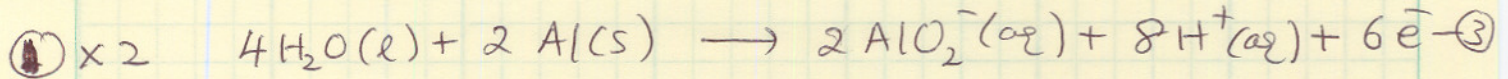
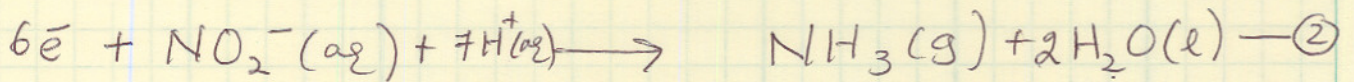
reduction rⁿ

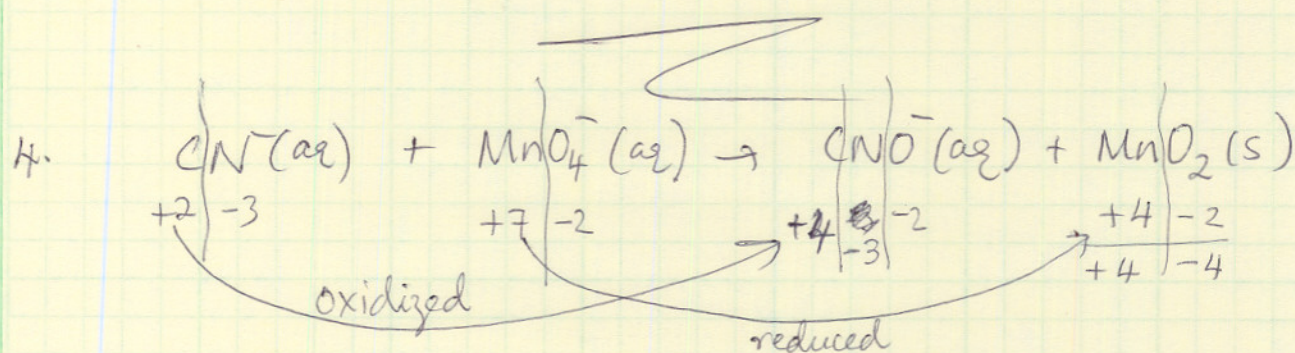
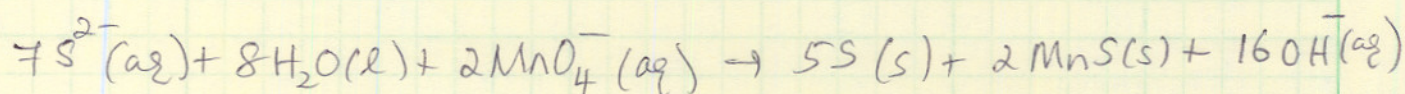
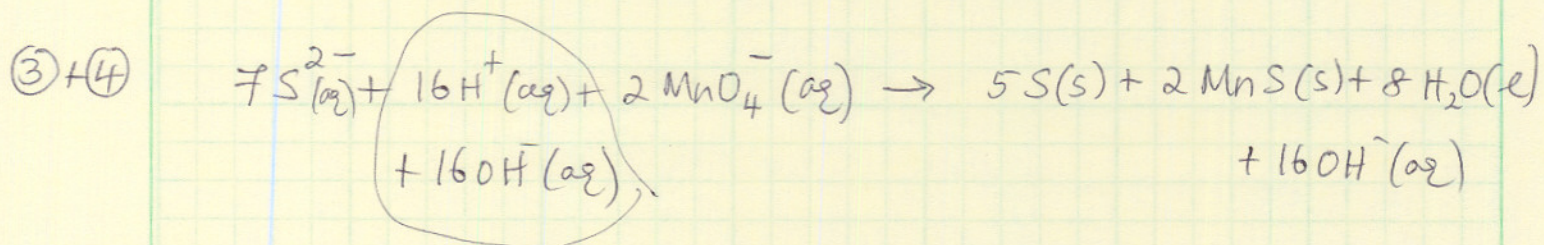
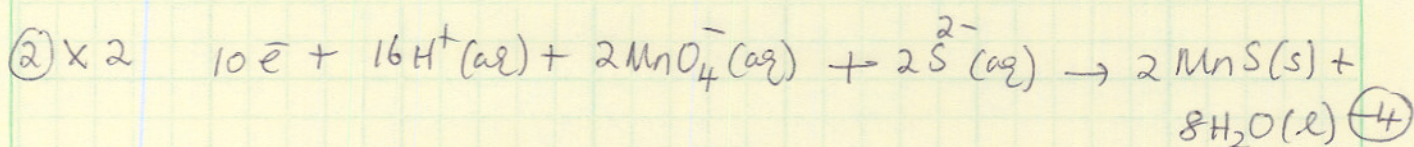
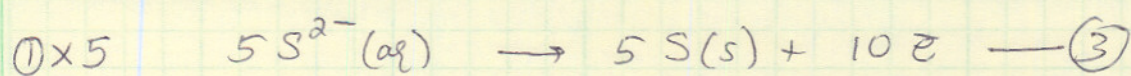


oxidation rⁿ

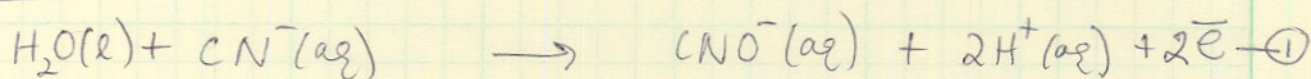


reduction r^n

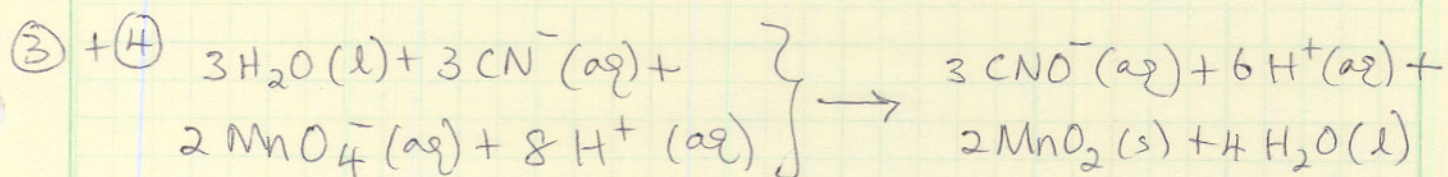
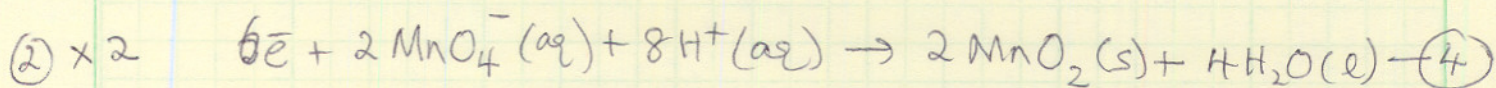
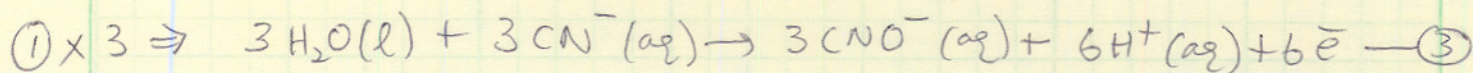
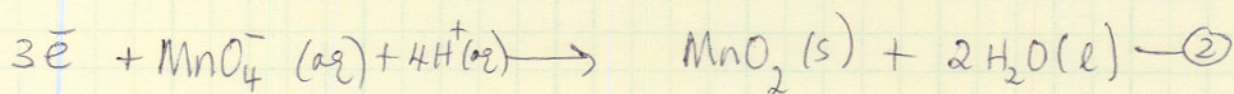




Oxidation rⁿ



reduction reaction



+ 2OH⁺

+ H₂O(l)

+ 2OH⁻(aq)

