

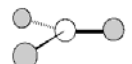
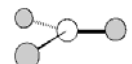
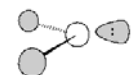

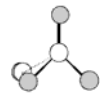
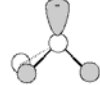

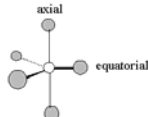

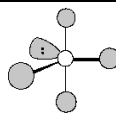
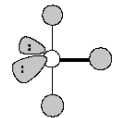
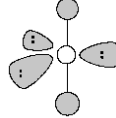
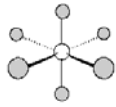
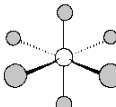
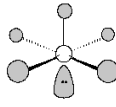
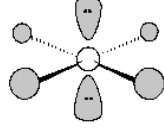


Molecular Geometry using the VSEPR model

# e pairs around central atom	Electron pair geometry	σ bonds	Lone pairs	Molecular shape	Bond angles
2	 linear	2	0	 linear	180°
3	 trigonal planar	3	0	 trigonal planar	120°
		2	1	 bent or angular	Less than 120°
4	 tetrahedral	4	0	 tetrahedral	109.5°
		3	1	 pyramidal	Less than 109.5°
		2	2	 bent or angular	Less than 109.5°
5	 trigonal bipyramidal	5	0	 trigonal bipyramidal	120° and 90°
		4	1	 see-saw or pseudo-tetrahedral	Less than 120° & 90°
		3	2	 T shape	90°
		2	3	 linear	180°
6	 octahedral	6	0	 octahedral	90°
		5	1	 square pyramidal	90°
		4	2	 square planar	90°